

Download File PDF Chemistry Workbook Chapter 15 Water And Aqueous Systems Answers

#Jenny



Finally I get this ebook, thanks for all these I can get now!

#Rio



Cool! I'am really happy

#Markus Jensen



I did not think that this would work, my best friend showed me this website, and it does! I get my most wanted eBook

#Hun Tsu



wtf this great ebook for free?!

#Che Salsa



My friends are so mad that they do not know how I have all the high quality ebook which they do not!

#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

Section 15.1
Water and its Properties

OBJECTIVES:

- Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.
- Describe the structure of ice.

The Water Molecule: a Review

- Water is a simple tri-atomic molecule, H_2O .
- Each O-H bond is _____, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values).
- Bond angle of water = 105° .
- due to the _____, the O-H bond polarities do not cancel. This means _____.

- Water's **heat capacity** and **ability to hydrogen bond** gives it many special properties!
- Water molecules are attracted to one another by dipole interactions.
- This hydrogen bonding gives water:
 - its _____ and
 - its _____.

a) High Surface Tension?

- liquid water acts like it has a "skin"
 - _____ glass of water bulges over the top
- Water forms round drops
 - _____ spray water on grassy surface
- All because water hydrogen bonds.

Surface Tension

- One water molecule can hydrogen bond to another because of this electrostatic attraction.
- Also, hydrogen bonding occurs with many other molecules surrounding them on all sides.

A water molecule in the middle of a solution is pulled in all directions.

[Download PDF version of :](#)
Chemistry Workbook Chapter 15 Water And Aqueous Systems Answers