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Chemistry 11 KEY
Unit 9 Review

7. Using the Energy Level Diagram for Many Electron Atoms (p.153 of SW), give the electron configuration for each of the following ions (You may use core notation.)

$P^{3-} (44) [Ar] 3s^2 3p^6$
 $Mg^{2+} (40) [Ne] 2s^2 2p^6$
 $Br (35) [Ar] 4s^2 3d^5 4p^6$
 $As^{3-} (33) [Ar] 4s^2 3d^5 4p^6$
 $Ti^{2+} (22) [Ar] 3d^2 4s^0 5p^0$
 $Li^+ (3) [He] 1s^2$

8. What is the letter which represents the principal quantum number? n
What does it stand for? energy level

9. The elements Co-Lu are filling up 3 f orbitals.

10. In order to become stable,

an atom of Ca will <u>lose</u> <u>2</u> electrons and become the ion <u>Ca^{2+}</u>
an atom of Se will <u>gain</u> <u>2</u> electrons and become the ion <u>Se^{2-}</u>
an atom of K will <u>lose</u> <u>1</u> electron and become the ion <u>K^+</u>
an atom of Br will <u>gain</u> <u>1</u> electron and become the ion <u>Br^-</u>
an atom of N will <u>gain</u> <u>3</u> electrons and become the ion <u>N^{3-}</u>
an atom of As will <u>gain</u> <u>3</u> electrons and become the ion <u>As^{3-}</u>
an atom of Al will <u>lose</u> <u>3</u> electrons and become the ion <u>Al^{3+}</u>
an atom of Te will <u>gain</u> <u>2</u> electrons and become the ion <u>Te^{2-}</u>

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